

Amendments to the Specification (where added material is shown in underlined type, deleted material is shown in ~~strikeout type~~.)

Please replace the paragraph starting on page 1, line 4 with following rewritten paragraph:

CROSS REFERENCE TO RELATED APPLICATIONS

This application is a continued examination application of continuation-in-part application U.S. Serial No. 10/719,504, filed on November 21, 2003 (published as Publication No. 2004/0115492 on June 17, 2004), which is a continuation-in-part of ~~application applications~~ U.S. Serial No. 10/602,536 filed June 23, 2003 and U.S. Serial No. 10/184,264 filed June 27, 2002 (published as Publication No. 2003/0022035 on January 30, 2003). This application is related to and contains common subject matter with U.S. Serial No. 09/186,766 filed November 5, 1998; now U.S. Patent No. 6,187,465 issued February 13, 2001 (the '465 patent), which claims the benefit of U.S. provisional application Serial No. 60/064,692 filed November 7, 1997. This application is not a continuation-in-part of the latter application, U.S. Serial No. 09/186,766, as stated in the parent application, U.S. Serial No. 10/184,264.

Please replace the paragraph starting on page 4, line 11 with following rewritten paragraph:

The present process avoids the difficult path of attempting to strip and capture the carbon dioxide from stack gases and without attempting to carry out separate chemical reactions of carbon dioxide to attempt to produce useful products. The process and system of the present invention uses unique gasification technology combined with fuel cells to generate electricity at high efficiency. This is accomplished by taking advantage of a very unique property of fuel cells - namely, the two anodic and cathodic reactions are separated by an electronically conducting membrane that keeps the product gases separate. In this way, a combustible feed gas can be fully oxidized in the first half-cell of the fuel cell without being commingled with the final products of the air reduction in the second half-cell electrode, i.e., N₂. For example, in coal

gasification, synthesis gas is formed consisting predominantly of hydrogen and carbon monoxide. This synthesis gas is fed into the first half-cell, i.e., the anode or negative terminal side, of the fuel cell, such as the solid oxide or molten carbonate types, where it is oxidized to water and carbon dioxide. These gases are not diluted by the typical nitrogen remaining after oxygen reduction in the second or remaining half-cell, i.e., the cathode side or positive terminal, of the fuel side. Nitrogen and combustion gases are commingled when combustion air is used in boilers or furnaces. Thus, in the fuel cell, the synthesis gas (syngas) is oxidized without being combusted with air and without being diluted by other gases. The fuel cell-produced water and carbon dioxide are simply separated from each other by condensing the liquid water and allowing the carbon dioxide to return to the gasifier. The carbon dioxide being injected into the high temperature gasifier undergoes a reaction with the high temperature carbonaceous feed to form more carbon monoxide, repeating the cycle.

Please replace the paragraph starting on page 5, line 3 with following rewritten paragraph:

By means of the present process and system, the carbon dioxide in the fuel cell is easily kept separate from the air side and any nitrogen. This carbon dioxide can be recycled back to the gasifier in nearly pure form. Likewise steam in pure form can be recycled as well in different amounts under gasifier control system requirements to maintain the ideal hydrogen to carbon monoxide ratio in the range of about 1.2 to about 1.75. This steam helps maintain a high hydrogen content in the gasifier so that a portion of the gasifier-produced syngas can be used downstream in a chemical reactor such as a Fischer-Tropsch reaction system for the production of a variety of useful chemicals ranging from methanol to paraffin waxes. These in turn are used to make useful chemicals such as naphtha, gas oil, and kerosine, or agricultural chemicals or carbide abrasives. The latter are not ever burned in their lifecycle, and they sequester the carbon forever. Thus, the carbon monoxide is used to produce useful chemicals instead of discarding the valuable carbon source in the carbon dioxide. The carbon balance of the plant is maintained such that the mass of carbon input in the waste feed is equal to the carbon mass leaving the plant as valuable hydrocarbon products; not carbon dioxide.

Please replace the paragraph starting on page 5, line 19 with the following rewritten paragraph:

What has been achieved is a chemical plant merged with a power plant that produces useful hydrocarbon products, high efficiency electric power without substantially any carbon dioxide or other greenhouse gas emissions. And, most importantly gasification is much more flexible than a refinery or a coal boiler, since a wide variety of waste streams can be used as the feed material. Thus, this solves two serious problems.

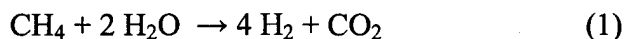
Please replace the paragraph starting on page 8, line 19 with the following rewritten paragraph:

The oxidized syngas (consisting essentially of hydrogen and carbon monoxide) leaves anode 42 of fuel cell 26 mostly as water vapor and carbon dioxide. This stream of oxidized syngas passes via line 48 into air-cooled condenser 50, where the water vapor is condensed into liquid water and is removed from the condenser bottoms via line 52 for reuse. Wastewater recovered from a municipal sewage system can be used in gasifier 12. However, all or a portion of the relatively pure water in line 52 can be sold or recycled and combined with the wastewater passing into gasifier 12 via line 38. The carbon dioxide gas is not condensed in condenser 50 and passes through into the condenser overhead as carbon dioxide gas to be fed back to the gasifier 12 via line 36. The carbon dioxide in high temperature gasifier 12 reacts therein with the carbonaceous feed material to form more syngas to further assist in the overall reaction. CO₂ or other greenhouse gases can be passed into gasifier 12 via line 56 and line 34 to maintain the desired H/C ratio of the feedstock.

Please replace the paragraph starting on page 9, line 31 with the following rewritten paragraph as follows:

The plots shown in FIGS. 2-10 are based on calculations performed by the method of the Gibbs Free Energy Minimization to yield gas compositions at thermodynamic equilibrium from

the lowest temperature of 200°C up to 2000°C. The chemistry is started by placing methane (CH₄) and steam (H₂O) at one atmosphere in the gaseous state (subscript, g) in a vessel at 200°C. After waiting a sufficient amount of time, the compounds react slightly and form a small quantity of hydrogen (H₂) and carbon dioxide (CO₂) as shown in Fig. 2. This composition of the gas mixture is that which occurs if the chemical kinetics were fast enough to allow the reaction to reach completion in the time allotted. The following two reactions are occurring simultaneously:



Please replace the paragraph starting on page 12, line 12 with the following rewritten paragraph as follows:

Unexpectedly, a much-preferred solution has now been discovered to optimize this fuel cell link that has been overlooked and not exploited previously. It involves using elevated steam feed and CO₂ simultaneously with complex waste streams that have higher carbon/hydrogen ratios than simpler compounds such as methane. This approach appears to be contrary to conventional wisdom and practice, which suggests that to achieve high ~~higher~~ hydrogen concentrations at high temperature, the worst option is to increase the carbon content of the feed. However, for carbon dioxide this simplistic logic has been found to be very wrong.

Please replace the paragraph starting on page 12, line 23 with the following rewritten paragraph:

Reaction (3) is already 40.5% 42.1% by volume hydrogen (i.e. mole percent), which is comparable to the hydrogen levels in FIGS. 3-6. Therein, one would have expected about 46% by volume H₂. Reaction (3) stoichiometry is the rough optimum, maximizing hydrogen content. Varying the stoichiometric quantities of the reactants produces less than optimum hydrogen. It is noteworthy that the addition of CO₂ to the feed reduces the requirements for steam below stoichiometric requirements. In fact, there is an optimum combination of using both CO₂ and steam.

Please replace the paragraph starting on page 13, line 13 with following rewritten paragraph:

One notes for a given carbonaceous feedstock with the empirical formula, $C_aH_bO_c$, one can adjust the amount of CO_2 , "D", to satisfy the Fischer-Tropsch synthesis requirements for the products desired.

Please replace the paragraph starting on page 13, line 16 with following rewritten paragraph:

To achieve higher hydrogen concentrations at high temperature to drive the fuel cells, increased feedstock hydrogen content together with an excess steam below stoichiometric levels, (5a-5c-D), is allowed and is combined with the recycled fuel cell carbon dioxide, D. As shown in Fig. 7-10, this provides the chemistry at thermodynamic equilibrium that achieves a high ~~higher~~ hydrogen-rich syngas that remains high and steady in hydrogen over a broad high temperature range up to and beyond 1300°C without catalysts.